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Do practice problems: When carbon disulfide is formed from its elements, heat is absorbed. Calculate the amount of heat (in kJ) absorbed when 5.66g of carbon disulfide is formed. $C(s) + 2S(s) \rightarrow CS_2(l)$ $\Delta H = 89.3 \text{ kJ}$. The production of iron and carbon dioxide from iron (III) oxide and carbon monoxide is an exothermic reaction.

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Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words ΔH is positive? I. $H_2O(l) + 10,5\text{kcal} \rightarrow H_2O(g)$ ΔH_1 II. $2NH_3 + 22\text{kcal}$

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CHAPTER 17, Thermochemistry (continued) GUIDED PRACTICE PROBLEM 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5°C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5°C

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Thermochemistry Practice Problems (Ch. 6) 1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 °C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21 °C? Explain your reasoning. 2.

Calorimetry Problems, Thermochemistry Practice, Specific ...

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